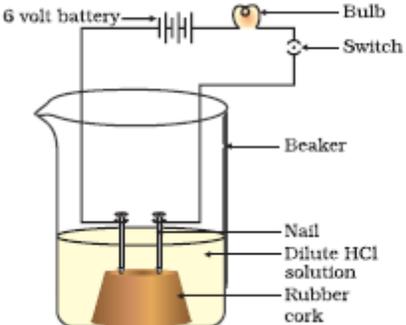


	(a) 0 (b) 3 (c) 7 (d) 9	
8.	<p>The diagram given here shows the relationship between bases and alkalis. Explain what does it mean?</p> <div style="text-align: center;"> </div>	1
9.	Distilled water is a poor conductor of electricity whereas rain water is a good conductor. Why?	1
10.	Why does an aqueous solution of acid conduct electricity?	1
11.	Name the sodium compound used to remove the permanent hardness of water.	1
12.	A white chemical compound becomes hard on mixing proper quantity of water. It is also used in surgery to maintain joints in a fixed position. Name the chemical compound.	1
13.	What is the commercial name of calcium sulphate hemi hydrate?	1
14.	What will happen if heating is not controlled while preparing plaster of Paris?	2
15.	<p>CO₂ gas is passed through lime water. Write the chemical equation to show the change.</p> <p>(a) What happens if excess of CO₂ is passed through lime water?</p> <p>(b) What is the product formed?</p> <p>(c) How will you represent the reaction?</p>	2
16.	A metal compound reacts with dilute hydrochloric acid. The gas evolved here extinguishes a burning candle. What is this metal compound? The other product formed here is MgCl ₂ . Write the chemical equation involved here.	2
17.	What is meant by 'hydrated' and 'anhydrous' salts? Explain with an example.	2
18.	What is the common name of Na ₂ CO ₃ .10 H ₂ O? Write any two uses of the compound.	2
19.	When the concentrated aqueous solution of substance X is electrolysed, then NaOH, Cl ₂ , and H ₂ are produced. Name the substance X. What is the special name of this process?	2
20.	Describe how sodium hydrogen carbonate is produced on large scale. Write the equation of the reaction involved.	2

21.	<p>A white powdery substance having strong smell of chlorine is used for disinfecting water. Identify the substance. Give its chemical name and write the chemical equation for its preparation.</p>	2
22.	<div style="text-align: center;">  </div> <p>Will the bulb glow if glucose solution is taken in the beaker instead of hydrochloric acid? Why?</p>	2
23.	<p>Write the name and formula of one salt each which contains:</p> <p>(a) two molecules of water of crystallisation.</p> <p>(b) five molecules of water of crystallisation.</p> <p>(c) ten molecules of water of crystallisation.</p>	3
24.	<p>Describe how washing soda is produced starting from sodium chloride. Write the equation of all reactions involved.</p>	3
25.	<p>Complete the following chemical reactions:</p> <p>(a) $\text{NaOH} + \text{Zn} \xrightarrow{\text{heat}} \text{-----} + \text{-----}$</p> <p>(b) $\text{NaHCO}_3 \rightarrow \text{-----} + \text{-----} + \text{-----}$</p> <p>(c) $\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{-----} + \text{-----} + \text{-----}$</p>	3