



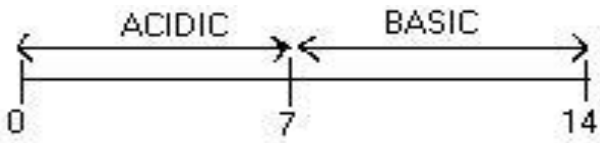
INDIAN SCHOOL DARSAIT
DEPARTMENT OF CHEMISTRY
WORKSHEET

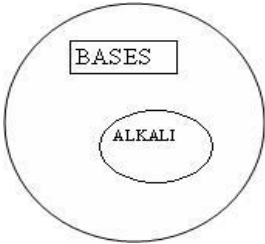


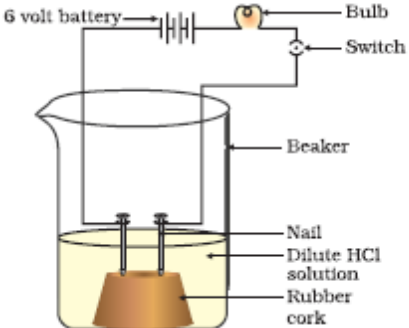
Subject : CHEMISTRY Chapter : Acids, Bases and Salts Date of Worksheet : 20-5-18

Resource Person: Mr. Harikrishnan P Date of submission : 03-6-18

Name of the Student : _____ Class & Division : X---- Roll Number : -----

1.	What is an indicator? Name an indicator derived from lichens.	1
2.	A farmer has found that pH of soil in his fields is 4.2. Name any two chemical material that he can mix with the soil to adjust the pH.	1
3.	Two solutions A and B have pH values of 3.0 and 9.5 respectively. Which of these will turn litmus solution from blue to red and which will turn phenolphthalein from colourless to pink?	1
4.	In which of the following cases will a gas be given off? (a) Hydrochloric acid is added to sodium carbonate. (b) Sulphuric acid is added to sodium hydroxide. (c) Sulphuric acid is added to copper oxide. (d) Hydrochloric acid is added to zinc oxide.	1
5.	What are the ions present in the solutions of following substances? (a) Hydrochloric acid (b) Nitric acid (c) Sulphuric acid (d) Potassium hydroxide (e) Magnesium hydroxide	1
6.	Two solutions X and Y are tested with universal indicator. Solution X turns orange whereas solution Y turns red. Which of the solutions is a stronger acid?	1
7.	The pH of a substance is a measure of its acidity as shown in the scale below:  When Ananya complained of acidity, her doctor prescribed a liquid antacid which gave her relief. Which of these could be the pH of the antacid?	1

	(a) 0 (b) 3 (c) 7 (d) 9	
8.	<p>The diagram given here shows the relationship between bases and alkalis. Explain what does it mean?</p> 	1
9.	Distilled water is a poor conductor of electricity whereas rain water is a good conductor. Why?	1
10.	Why does an aqueous solution of acid conduct electricity?	1
11.	Name the sodium compound used to remove the permanent hardness of water.	1
12.	A white chemical compound becomes hard on mixing proper quantity of water. It is also used in surgery to maintain joints in a fixed position. Name the chemical compound.	1
13.	What is the commercial name of calcium sulphate hemi hydrate?	1
14.	What will happen if heating is not controlled while preparing plaster of Paris?	2
15.	<p>CO₂ gas is passed through lime water. Write the chemical equation to show the change.</p> <p>(a) What happens if excess of CO₂ is passed through lime water?</p> <p>(b) What is the product formed?</p> <p>(c) How will you represent the reaction?</p>	2
16.	A metal compound reacts with dilute hydrochloric acid. The gas evolved here extinguishes a burning candle. What is this metal compound? The other product formed here is MgCl ₂ . Write the chemical equation involved here.	2
17.	What is meant by 'hydrated' and 'anhydrous' salts? Explain with an example.	2
18.	What is the common name of Na ₂ CO ₃ .10 H ₂ O? Write any two uses of the compound.	2
19.	When the concentrated aqueous solution of substance X is electrolysed, then NaOH, Cl ₂ , and H ₂ are produced. Name the substance X. What is the special name of this process?	2
20.	Describe how sodium hydrogen carbonate is produced on large scale. Write the equation of the reaction involved.	2

21.	<p>A white powdery substance having strong smell of chlorine is used for disinfecting water. Identify the substance. Give its chemical name and write the chemical equation for its preparation.</p>	2
22.	<div style="text-align: center;">  </div> <p>Will the bulb glow if glucose solution is taken in the beaker instead of hydrochloric acid? Why?</p>	2
23.	<p>Write the name and formula of one salt each which contains:</p> <p>(a) two molecules of water of crystallisation.</p> <p>(b) five molecules of water of crystallisation.</p> <p>(c) ten molecules of water of crystallisation.</p>	3
24.	<p>Describe how washing soda is produced starting from sodium chloride. Write the equation of all reactions involved.</p>	3
25.	<p>Complete the following chemical reactions:</p> <p>(a) $\text{NaOH} + \text{Zn} \xrightarrow{\text{heat}}$ ----- + -----</p> <p>(b) $\text{NaHCO}_3 \rightarrow$ ----- + ----- + -----</p> <p>(c) $\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \rightarrow$ ----- + ----- + -----</p>	3